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Model 2 – Prefixes and Suffixes

Prefix	Numerical Value	Molecular Formula	Name of Compound
mono-	1	$\text{BCl}_3$	Boron trichloride
di-	2	$\text{SF}_6$	Sulfur hexafluoride
tri-	3	$\text{IF}_7$	Iodine heptafluoride
tetra-	4	$\text{NI}_3$	Nitrogen triiodide
penta-	5	$\text{N}_2\text{O}_4$	Dinitrogen tetroxide
hexa-	6	$\text{Cl}_2\text{O}$	Dichlorine monoxide
hepta-	7	$\text{P}_4\text{O}_{10}$	Tetraphosphorus decoxide
octa-	8	$\text{B}_2\text{H}_6$	Diboron hexahydride
nona-	9	$\text{Br}_2\text{O}_9$	Dibromine nonoxide
deca-	10	$\text{ClF}$	Chlorine monofluoride

5. Examine the prefixes in Model 2. Fill in the numerical value that corresponds to each prefix.  
 6. What suffix (ending) do all the compound names in Model 2 have in common?

*They all end in "ide".*

7. Carefully examine the names of the compounds in Model 2. When is a prefix NOT used in front of the name of an element?  
*When the first element is one. (Examples:  $\text{IF}_7$  - iodine heptafluoride,  $\text{NI}_3$  - nitrogen triiodide)*

8. Consider the compound NO.

- a. Which element, nitrogen or oxygen, would require a prefix in the molecule name? Explain your answer.

*Oxygen would require a prefix of "mono" because it is the second element and there is one oxygen. Nitrogen does not require a prefix because it is the first element and there is one nitrogen.*

- b. Name the molecule NO.

*Nitrogen monoxide (Notice, the "o" of monoxide dropped) → not monoxide*

9. Find two compounds in Model 2 that contain a subscript of "4" in their molecular formula.

- a. List the formulas and names for the two compounds.

*$\text{N}_2\text{O}_4$  - dinitrogen tetroxide ||  $\text{P}_4\text{O}_{10}$  - Tetraphosphorus decoxide*

- b. What is different about the spelling of the prefix meaning "four" in these two names?

*The "a" has been dropped from the end of the prefix tetra.*

Key

Formula	Name
$\text{Ag}_2\text{S}$	Silver sulfide
$\text{Al}_2\text{O}_3$	Aluminum oxide
$\text{SO}_2$	Sulfur dioxide
$\text{Fe}(\text{OH})_2$	Iron (II) hydroxide
$\text{CrCl}_3$	Chromium (III) chloride
$\text{CaF}_2$	Calcium fluoride
$\text{PCl}_5$	Phosphorus pentachloride
$\text{XeF}_4$	Xenon tetrafluoride
$\text{NH}_4\text{SO}_4$	Ammonium sulfate
$\text{Hg}_2(\text{NO}_3)_2$	Mercury (I) nitrate
$\text{NiBr}_2$	Nickel (II) bromide
$\text{SF}_6$	Sulfur hexafluoride
$\text{CuCl}$	Copper (I) chloride
$\text{NaC}_2\text{H}_3\text{O}_2$	Sodium acetate
$\text{AsI}_3$	Arsenic triiodide
$\text{OF}_2$	Oxygen difluoride
$\text{HgSO}_4$	Mercury (II) sulfate
$\text{H}_2\text{CO}_3$	Carbonic acid
$\text{Fe}_2(\text{PO}_4)_3$	Iron (III) phosphate
$\text{K}_2\text{CrO}_4$	Potassium chromate
$\text{H}_2\text{S}$	Hydrogen sulfide
$\text{NO}$	Nitrogen monoxide
$\text{CuClO}_2$	Copper (I) chlorate
$\text{Zn}(\text{OH})_2$	Zinc hydroxide
$\text{AgC}_2\text{H}_3\text{O}_2$	Silver acetate
$\text{TiCl}_4$	Titanium (IV) chloride
$\text{BaS}$	Barium sulfide
$\text{Pb}_3(\text{PO}_4)_2$	Lead (II) phosphate
$\text{CrClO}_4$	Cesium perchlorate
$\text{XeCl}_4$	Xenon tetrachloride
$\text{Fe}(\text{NO}_3)_3$	Iron (III) nitrate
$\text{Fe}(\text{NO}_2)_3$	Iron (III) nitrite
$\text{FeN}$	Iron (III) nitride
$\text{Ni}(\text{OH})_2$	Nickel (II) hydroxide
$\text{Mn}(\text{OH})_2$	Manganese (II) hydroxide
$\text{KClO}$	Potassium hypochlorite
$\text{KClO}_2$	Potassium chlorite
$\text{KClO}_3$	Potassium chlorate
$\text{KClO}_4$	Potassium perchlorate
$\text{Cl}_2$	Chlorine

Homework. Name each compound. (These are a mixed selection of the five types above)

- SILVER SULFIDE  $\text{Ag}_2\text{S}$
- ALUMINUM OXIDE  $\text{Al}_2\text{O}_3$
- SULFUR DIOXIDE  $\text{SO}_2$
- IRON (II) HYDROXIDE  $\text{Fe}(\text{OH})_2$
- CHROMIUM (III) CHLORIDE  $\text{CrCl}_3$
- CALCIUM FLUORIDE  $\text{CaF}_2$
- PHOSPHORUS PENTACHLORIDE  $\text{PCl}_5$
- XENON TETRAFLUORIDE  $\text{XeF}_4$
- AMMONIUM SULFATE  $(\text{NH}_4)_2\text{SO}_4$
- MERCURY (I) NITRATE  $\text{Hg}_2(\text{NO}_3)_2$
- NICKEL (II) BROMIDE  $\text{NiBr}_2$
- SULFUR HEXAFLUORIDE  $\text{SF}_6$
- COPPER (I) CHLORIDE  $\text{CuCl}$
- SODIUM ACETATE  $\text{NaC}_2\text{H}_3\text{O}_2$
- ARSENIC TRIIODIDE  $\text{AsI}_3$
- OXYGEN DIFLUORIDE  $\text{OF}_2$
- MERCURY (II) SULFATE  $\text{HgSO}_4$
- CARBONIC ACID  $\text{H}_2\text{CO}_3$
- IRON (III) PHOSPHATE  $\text{Fe}_2(\text{PO}_4)_3$
- POTASSIUM CHROMATE  $\text{K}_2\text{CrO}_4$
- HYDROGEN SULFIDE  $\text{H}_2\text{S}$
- DINITROGEN MONOXIDE  $\text{N}_2\text{O}$
- COPPER (I) CHLORATE  $\text{CuClO}_2$
- ZINC HYDROXIDE  $\text{Zn}(\text{OH})_2$
- SILVER ACETATE  $\text{AgC}_2\text{H}_3\text{O}_2$
- TITANIUM (IV) CHLORIDE  $\text{TiCl}_4$
- BARIUM SULFIDE  $\text{BaS}$
- LEAD (II) PHOSPHATE  $\text{Pb}_3(\text{PO}_4)_2$
- CESIUM PERCHLORATE  $\text{CsClO}_4$
- XENON TETRACHLORIDE  $\text{XeCl}_4$
- IRON (III) NITRATE  $\text{Fe}(\text{NO}_3)_3$
- IRON (III) NITRITE  $\text{Fe}(\text{NO}_2)_3$
- IRON (III) NITRIDE  $\text{FeN}$
- AMMONIUM HYDROXIDE  $\text{NH}_4\text{OH}$
- MANGANESE (II) HYDROXIDE  $\text{Mn}(\text{OH})_2$
- POTASSIUM HYPOCHLORITE  $\text{KClO}$
- POTASSIUM CHLORITE  $\text{KClO}_2$
- POTASSIUM CHLORATE  $\text{KClO}_3$
- POTASSIUM PERCHLORATE  $\text{KClO}_4$
- CHLORINE  $\text{Cl}_2$
- PERCHLORIC ACID  $\text{HClO}_4$
- TITANIUM (IV) CYANIDE  $\text{Ti}(\text{CN})_4$
- PHOSPHORUS TRICHLORIDE  $\text{PCl}_3$
- COPPER (II) NITRATE  $\text{Cu}(\text{NO}_3)_2$
- ALUMINUM PERCHLORATE  $\text{Al}(\text{ClO}_4)_3$
- SILVER NITRATE  $\text{AgNO}_3$
- GOLD (III) SULFIDE  $\text{Au}_2\text{S}_3$
- AMMONIUM CARBONATE  $(\text{NH}_4)_2\text{CO}_3$
- SODIUM HYDROXIDE  $\text{NaOH}$
- URANIUM (VI) FLUORIDE  $\text{UF}_6$
- MAGNESIUM IODIDE  $\text{MgI}_2$
- CESIUM NITRATE  $\text{CsNO}_3$
- DINITROGEN PENTOXIDE  $\text{N}_2\text{O}_5$
- COPPER (II) SULFATE  $\text{CuSO}_4$
- SELENIC ACID  $\text{H}_2\text{SeO}_4$
- TETRAPHOSPHORUS HEXAOXIDE  $\text{P}_4\text{O}_6$
- MANGANESE (IV) OXIDE  $\text{MnO}_2$
- NICKEL (II) PHOSPHATE  $\text{Ni}_3(\text{PO}_4)_2$
- SODIUM SULFIDE  $\text{Na}_2\text{S}$
- COPPER (II) CYANIDE  $\text{Cu}(\text{CN})_2$
- LEAD (II) CARBONATE  $\text{PbCO}_3$
- OXYGEN DIFLUORIDE  $\text{OF}_2$
- SODIUM CHLORIDE  $\text{NaCl}$
- OXALIC ACID  $\text{H}_2\text{C}_2\text{O}_4$
- POTASSIUM PHOSPHATE  $\text{K}_3\text{PO}_4$
- NITROUS ACID  $\text{HNO}_2$
- TELLURIUM DICHLORIDE  $\text{TeCl}_2$
- ARSENIC ACID  $\text{H}_3\text{AsO}_3$
- COPPER (II) PHOSPHITE  $\text{Cu}_3(\text{PO}_3)_2$
- DIPHOSPHORUS PENTAOXIDE  $\text{P}_2\text{O}_5$
- LEAD (IV) PERMANGANATE  $\text{Pb}(\text{MnO}_4)_2$
- TIN (II) CHROMATE  $\text{SnCrO}_4$
- GOLD (III) CHLORIDE  $\text{AuCl}_3$
- LEAD (II) CHROMATE  $\text{PbCrO}_4$
- SODIUM SULFIDE  $\text{Na}_2\text{S}$
- MAGNESIUM CARBONATE  $\text{MgCO}_3$
- URANIUM (VI) SULFATE  $\text{U}(\text{SO}_4)_3$
- SULFUR TRIOXIDE  $\text{SO}_3$
- SODIUM CYANIDE  $\text{NaCN}$
- SODIUM PEROXIDE  $\text{Na}_2\text{O}_2$

Table of common polyatomic ions

Ion name	Formula/charge
acetate	$\text{C}_2\text{H}_3\text{O}_2^{-1}$
ammonium	$\text{NH}_4^{+1}$
bicarbonate	$\text{HCO}_3^{-1}$
carbonate	$\text{CO}_3^{-2}$
chromate	$\text{CrO}_4^{-2}$
cyanide	$\text{CN}^{-1}$
hydroxide	$\text{OH}^{-1}$
nitrate	$\text{NO}_3^{-1}$
nitrite	$\text{NO}_2^{-1}$
permanganate	$\text{MnO}_4^{-1}$
phosphate	$\text{PO}_4^{-3}$
sulfate	$\text{SO}_4^{-2}$
sulfite	$\text{SO}_3^{-2}$

Things to keep in mind:  
 a) Charges written as "-2" or "-2-" are considered the same. It doesn't matter in which order you put the number and sign.  
 b) A charge of "+1" or "-1" may be written as "+ or -", respectively.  
 c) The bicarbonate ion is sometimes called the "hydrogen carbonate" ion in chemistry classes, but not in the real world.  
 d) The acetate ion is sometimes shown as  $\text{CH}_3\text{COO}^{-1}$ ; it means exactly the same thing as the formula shown here.  
 e) There are many, many other polyatomic ions out there. Make sure you memorize those that your instructor finds important!

## Ionic Chemical Formulas/PreAICE Chemistry

NAME: \_\_\_\_\_ PERIOD \_\_\_\_\_

Directions: Using the Chemical Formulas program sheet, fill in the table below. Use subscripts & superscripts. Remember to use roman numerals to show the valence of the transition metals in the compound name.

METAL CATION (+)	CATION FORMULA	NONMETAL ANION (-)	ANION FORMULA	COMPOUND FORMULA	COMPOUND NAME
sodium	Na+1	hydroxide	OH-	NaOH	sodium hydroxide
magnesium	Mg+2	phosphate	PO4-3	Mg3(PO4)2	magnesium phosphate
calcium	Ca+2	chloride	Cl-	CaCl2	calcium chloride
gold(I)	Au+1	sulfite	SO3-2	Au2SO3	gold(I) sulfite
copper(II)	Cu+2	hydroxide	OH-	Cu(OH)2	copper(II) hydroxide
Iron(III)	Fe+3	telluride	Te-2	Fe2Te3	iron(III) telluride
magnesium	Mg+2	bromide	Br-1	MgBr2	magnesium bromide
manganese(II)	Mn+2	sulfide	S-2	MnS	manganese(II) sulfide
boron	B+3	oxide	O-2	B2O3	boron oxide
potassium	K+	nitrate	NO3-	KNO3	potassium nitrate
barium	Ba+2	bromide	Br-	BaBr2	barium bromide
strontium	Sr+2	sulfate	SO4-2	SrSO4	strontium sulfate
iron(III)	Fe+3	oxide	O-2	Fe2O3	iron(III) oxide
silver(I)	Ag+1	chloride	Cl-	AgCl	silver(I) chloride
aluminum	Al+3	carbonate	CO3-2	Al2(CO3)3	aluminum carbonate
gold(II)	Au+2	nitride	N-3	Au3N2	gold(II) nitride
cesium	Cs+1	phosphate	PO4-3	Cs3PO4	cesium phosphate
ammonium	NH4+	sulfide	S-2	(NH4)2S	ammonium sulfide
zinc(II)	Zn+2	nitrite	NO2-	Zn(NO2)2	zinc(II) nitrite
chromium(I)	Cr+1	nitride	N-3	Cr3N	chromium(I) nitride
mercury(II)	Hg+2	oxide	O-2	HgO	mercury(II) oxide
sodium	Na+	sulfite	SO3-2	Na2SO3	sodium sulfite
magnesium	Mg+2	oxide	O-2	MgO	magnesium oxide

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